Criterion –I

1.1.1 Curriculum Planning and Implementation

TEACHING LESSON PLANS

2018-2019 TO 2022-2023

St. Ann's College for Women

GORANTLA - GUNTUR

LESSON PLAN

2022 - 2023

Name of the Lecturer.....K. VANAJA.....

Department.....BIOTECHNOLOGY.....

LESSON PLAN

Name of the Lecturer: K.VANAJA ..... Name of the Department : Biotechnology

Class ..... III - BBC ..... Subject : ..... Apiculture .....

Name of the Topic : Modern methods of extraction of Honey

Hours required :

Learning Objectives : Modern methods were developed to overcome the drawbacks of indigenous methods

* In this method first of all work is done to improve the texture of the hive

* The newton model & Langstroth hive are used

Previous knowledge to be reminded :

Honey extraction process

Topic Synopsis :

* The common appliances for modern methods for beekeeping include : Typical movable hive
  * queen excludes
  * Honey extractor
  * uncapping knife
  * other equipment

* It is an artificial movable hive which is made by wooden box based on bee space theory
* This typical movable hive contains 6 parts
  * stand
  * Bottom board
  * Brood chamber
  * supper
  * Inner cover
  * Top cover

Examples / Illustrations

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Additional Inputs

—

Teaching Aids used

* Blackboard
* ppt Presentation

References cited

* Mishra R.C. – Honey bee and their management
* Singh – Beekeeping in India

Student Activity planned after teaching

* Preparation of Honey Bee trays & maintenance

Activity planned outside the Class Room, if any

Any other activity

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K. Lavanya.
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: K.VANAJA Name of the Department Biotechnology

Class III-BBC Subject : Pearl Culture

Name of the Topic : Artificeal Production of Pearls.

Hours required :

Learning Objectives : * Culture of Pearl is complex but sensitive.

i. Collection of oysters
ii. Preparation of Graft tissue
iii. Preparation of Nucleus
iv. Implantation
v. Rearing of oysters
vi. Harvesting

Previous knowledge to be reminded :

cultivation of pearl production,

Topic Synopsis : *the culture of Pearl is a complex but sensitive process. It involves the following steps:-

i. Collection of oysters
ii. Preparation of Graft tissue
iii. Preparation of Nucleus
iv. Implantation
v. Rearing of Oysters
vi. Harvesting.

Fresh water Pearl culture:- These are produced in china, Japan and United States. * Pinctada anomioides
* Pinctada Vulgaris etc.

Marine water Pearl culture:- These oysters only produce one Pearl and that takes about 2 years:- * Pinctada fucata
* Pinctada chemnitzi

Examples / Illustrations

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Additional Inputs

—

Teaching Aids used

* Black chart
* ppt presentation

References cited

*Bardach – Aquaculture farming and husbandry.

Student Activity planned after teaching

* Culture of microorganism for pond maintenance.

Activity planned outside the Class Room, if any

Collection of pearl oysters

Any other activity

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K. Vanaja
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: K.VANAJA Name of the Department : Biotechnology

Class III-BBC Subject : Pearl Culture

Name of the Topic : Chemical composition & Natural process of Pearl formation

Hours required :

Learning Objectives : * Chemical composition of Pearl is:-

82-86% - $CaCO_3$

10-14% - Conchiolin &

2-4% - water.

* Conchiolin & $CaCO_3$ are referred to as Nacre.

Previous knowledge to be reminded :

chemical comporition of pearl

Topic Synopsis : * The chemical composition of Pearl is:-

82-86% - $CaCO_3$

10-14% - conchiolin &

2-4% - water

* Together the Conchiolin & $CaCO_3$ are referred to as the Nacre.

* The natural Pearl oysters produce natural pearl as an adoptation against foreign materials.

* Foreign substance, for example a grain of sand (or) small insect must enter the shell of molluscs.

* The Mollusc begin to secrete Nacre against a grain of sand result in a smooth surface is obtained. This is called pearl, this pearl become a large pearl.

* The pearl is the result of self-defense of the Oyster.

Examples / Illustrations

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Additional Inputs

—

Teaching Aids used

* Blackboard
* Chart

References cited

* Haws Maria - Basics of Pearl farming.
* Alexander E. Farn - Pearls

Student Activity planned after teaching

* Graft tissue preparation
* Implantation techniques.

Activity planned outside the Class Room, if any

collection of pearl oyster -

Any other activity

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K. Vanaja
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: K.VANAJA Name of the Department : Biotechnology

Class II-BBC Subject Immunology & r-DNA

Name of the Topic : Bioinformatics, Database - Nucleotide & Protein

Hours required :

Learning Objectives : Definition

* data bases :- DNA Sequence database
Protein sequence databases
RNA sequence database

Previous knowledge to be reminded :

* Bio - meaning living
* Informatics meaning information science
* DNA, RNA & proteins structures.

Topic Synopsis : * Bioinformatics is information technology dealing with the maintenance and use of data in Molecular Biology using computers.

* Bioinformatics involves the collection, maintenance, distribution, analyse and usage of the large amount of data generated in Molecular Biology for Biological investigations.
* BLAST, FASTA are used to analyze the DNA Sequence.
* PIR, Swiss port are used to analyse the Protein sequence.

Examples / Illustrations

Sequence Analysis.

Additional Inputs

visit different web sit - biological dada sites

Teaching Aids used

* Black board
* Internet
* Biotechnology - V. kumaresen

References cited

* Biotechnology - V. kumaresan

Student Activity planned after teaching

* Collection of database.
* Observation of different databases.

Activity planned outside the Class Room, if any

Any other activity

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K. Vanaja
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: K.VANAJA Name of the Department: Biotechnology

Class: I-BBC Subject: Biomolecules & Analytics

Name of the Topic : Biological role of Vitamins, Bioenergetics

Hours required :

Learning Objectives : Biological role and deficiency manifestation of vitamin A, B, C, D and K.

* Study of energy.

Previous knowledge to be reminded : * Minerals & Vitamins are essential Nutrients - because acting in convert.

* Study of energy in living system.

* Energy is a capacity (or) ability to do work.

Topic Synopsis : * Vitamins may be regarded as organic compounds required in the diet in small amount to perform specific biological functions for normal maentenance of growth and health of the organisms.

* Funk coined the term vitamin in 1913.

* They are classified into: 1. Fat soluble (A, D, E, K)
2. Water Soluble (C & B group)

* Bioenergetics are related to initial & final state of energy of reactions.

* Entropy- order of randomness that is higher than the molecules than higher the Entropy.

* Enthalpy:- It is the heat content of the reacting system. All Enthalpy Changes.

Examples / Illustrations

Different Natural Sources of vitamins

Additional Inputs

Awareness programme on vitamins

Teaching Aids used

*Black board
*ppt presentation

References cited *Biochemistry - V. Sathyanarayanan

Student Activity planned after teaching

*Preparation of Vitamins chart.

Activity planned outside the Class Room, if any

*Multivitamin tablets.
able to list & explain vitamins essential to the health functioning of the human body.

Any other activity

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K. Uma Raja
Signature of the Lecturer

St. Ann's College for Women

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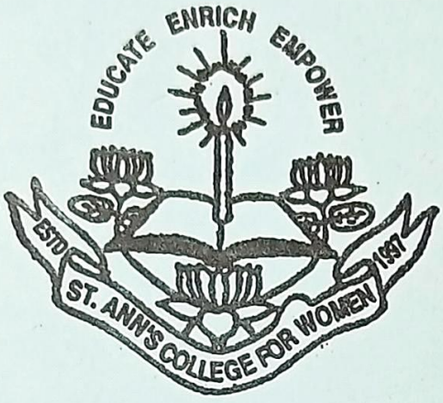

LESSON PLAN

2021 - 2022

Name of the Lecturer..... Ch. Rama Rao .....

Department..... Physics .....

UNIT-II Electromagnetic Induction

LESSON PLAN

Name of the Lecturer: Ch. Rama Rao Name of the Department : Physics

Class III B.Sc Subject : V

Name of the Topic : Electromagnetic Induction Introduction

Hours required : 3h

Learning Objectives :

* Concept of Faraday's laws
* Lenz's law
* Self Inductance and mutual Inductance.
* Self Inductance of a long solenoid.

Previous knowledge to be reminded :

Intermediate Syllabus

Topic Synopsis :

$$e = -\left(\frac{d\phi_B}{dt}\right)$$

Lenz's law: The direction of induced emf in a closed circuit is such that it opposes the original cause that produced it.

$$i = -\frac{1}{R}\left(\frac{d\phi_B}{dt}\right)$$

Magnetic flux $\int E \cdot dl = -\frac{d}{dt}\int B \cdot ds$

$$\text{curl } E = -\left(\frac{\partial B}{\partial t}\right)$$

$$e = -Blv \quad \text{(Induced emf)}$$

Self Inductance of a solenoid is

$$L = \mu_0 n^2 A l$$

Examples / Illustrations

Spark is observed in electrical switch is on (or) off quickly

Additional Inputs

notes

Teaching Aids used

Black - Board

References cited

unified text book.

Student Activity planned after teaching

Students are asked to practice the derivation

Activity planned outside the Class Room, if any

Students are asked to go through Intermediate Syllabus.

Any other activity

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Signature of the Lecturer

MODEREN PHYSICS.

LESSON PLAN

Name of theLecturer: Ch. Rama Rao Name of the Department : physics

Class III B.Sc Subject : VI

Name of the Topic : Atomic and molecular physics.

Hours required : 3h

Learning Objectives :

* Scattering of light
* Raman effect
* Raman shift
* applications.

Previous knowledge to be reminded :

Scattering concept in intermediate syllabus.

Topic Synopsis :

Raman effect:

When light incident on solid (or) liquid (or) gas it is scattered. The scattered light consists of greater frequency lines and lesser frequency lines is called Raman effect.

* greater frequency lines are called anti stokes lines
* lesser freq lines are called stokes lines.

Examples / Illustrations * Raman effect is pure molecular phenomena
* To Study the Structure of molecules.
* To study the Single, double and triple bonds etc.

Additional Inputs

notes

Teaching Aids used Black-Board.

References cited unified text book

Student Activity planned after teaching

students are asked questions on topic.

Activity planned outside the Class Room, if any

students are asked to practice the expt. study at home

Any other activity

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matter waves

LESSON PLAN

Name of theLecturer: Ch. Rama Rao. Name of the Department : Physics

Class III B.Sc Subject : VI

Name of the Topic : wave and particle duality of Radiation

Hours required : 3h

Learning Objectives :

* Dual nature of radiation
* de-Broglie theory, properties of matter
* Davisson-Germer electron diffraction expt. for wave nature supported.

Previous knowledge to be reminded :

Inter Syllabus.

Topic Synopsis : de-Broglie wave length

$$\lambda = \frac{h}{mv} = \frac{h}{\sqrt{2mE}} = \frac{h}{\sqrt{2m \cdot eV}} = \frac{12.26}{\sqrt{V}}$$

where h is planck's constant = 6.62×10^{-34} J-sec.

* phase velocity, wave velocity

$$v_p = \frac{\omega}{k} = \frac{c^2}{v}$$

* Explanation of electron diffraction experiment.

Examples / Illustrations

* The velocity of matter wave depend on the velocity of matter particle.

* The velocity of matter wave is greater than the velocity of light.

Additional Inputs

notes

Teaching Aids used

Black - Board

References cited

unified text book

Student Activity planned after teaching

practice question - general electron diffraction diagram.

Activity planned outside the Class Room, if any

Any other activity

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Signature of the Lecturer

Structure of Nucleus

LESSON PLAN

Name of theLecturer: Ch. Rama Rao Name of the Department : physics

Class III B.Sc Subject : VI

Name of the Topic : Introduction on nuclear physics. properties of nucleus

Hours required : 3h

Learning Objectives :

To Study the Structure of nucleus and properties of nucleus.

* Binding Energy $E = \Delta m \cdot c^2$

Previous knowledge to be reminded :

Intermediate Concepts.

Topic Synopsis :

* nuclear mass : $(Zm_p + Nm_n) - m_{nucleus}$
* nuclear radius $R \propto A^{1/3}$.
  $R = 1.1 A^{1/3}$.
* nuclear charge
* nuclear spin, quantum states
* Electric quadrupole moment.

Examples / Illustrations

* Scattering of α-particles, Rutherford concluded that the atom of any element consists of central core called nucleus.

Additional Inputs

velocity

Teaching Aids used

Black-Board.

References cited

Prescribed text book

Student Activity planned after teaching

Solve problems on nuclear radius and B.E.

Activity planned outside the Class Room, if any

Solve some more problems on the above topics.

Any other activity

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LESSON PLAN

Name of theLecturer: Ch. Rama Rao Name of the Department : physics

Class III B.Sc Subject : VI

Name of the Topic : β-decay

Hours required : 2h

Learning Objectives :

* To study about neutrino Hypothesis.
* Gamows explanation on β-decay.
* Fermi's theory of β-decay

Previous knowledge to be reminded :

syllabus covered in last class

Topic Synopsis :

According to yukawa's theory

$$n \rightarrow p + \pi^- \rightarrow p + e^- + \bar{\nu}$$

where e^- is electron and $\bar{\nu}$ is antineutrino.

neutron $\rightarrow {}_1H^1$ (proton) + electron + neutrino.

$$_0n^1 \rightarrow {}_1H^1 + {}_{-1}e^0 + \nu$$

Mass $1 \rightarrow 1 + 0 + 0$

Charge $0 \rightarrow 1 + -1 + 0$.

Examples / Illustrations —

Additional Inputs Notes

Teaching Aids used Black - Board.

References cited unified text book.

Student Activity planned after teaching —

Activity planned outside the Class Room, if any

students are asked to write self test on above topic.

Any other activity —

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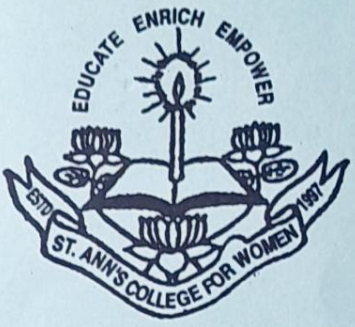


LESSON PLAN

2020 - 2021

Name of the Lecturer..... B. JOYCE .....

Department..... CHEMISTRY .....

LESSON PLAN

Name of theLecturer: B. Joyce Name of the Department : Chemistry

Class : III MBC, MPC, BBC Subject : Coordination Chemistry

Name of the Topic : Stability of Complexes

Hours required : 1 hr

Learning Objectives : To learn about.

1. Factors affecting Stability of complexes
2. Types of stabilities

Previous knowledge to be reminded :

Dissassociation, association
Complexes, coordination

Topic Synopsis :

Constant or formation Constant.

→ The reciprocal of β [$\frac{1}{\beta}$] is called instability constant or dissociation constant

→ The complex forming reversible reaction [$M^{n+} + nL \rightleftharpoons ML_n$] does not take place in a single step. It takes place in Several Steps

→ The steps involved in the complex [ML_n] formation are represent as.

$M + L \rightleftharpoons ML$ (charges omitted)

$ML + L \rightleftharpoons ML_2$

$ML_2 + L \rightleftharpoons ML_3$

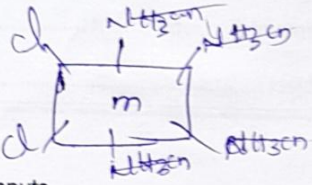
$ML_3 + L \rightleftharpoons ML_4$

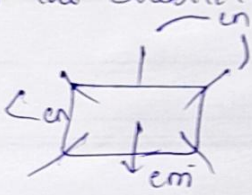
$\vdots \quad \vdots \quad \vdots$

$ML_{n-1} + L \rightleftharpoons ML_n$

Examples / Illustrations

Ring formations - more the no. of rings more is the stability





3-membered ring

Additional Inputs

The stability of metal complex is governed by two different aspects such as thermodynamic and kinetic stabilities.

Teaching Aids used

PPT - youtube videos.

References cited

Text book on
Coordination chemistry - PR publications

Student Activity planned after teaching

Excercise with Examples
slip test, Assignment

Activity planned outside the Class Room, if any

Google search on stability of different types of complexes.

Any other activity

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LESSON PLAN

Name of theLecturer: B. JOYCE Name of the Department : CHEMISTRY

Class : III BBC, MBC, & MPC Subject : Coordination Chemistry

Name of the Topic : THERMODYNAMIC AND KINETIC STABILITY.

Hours required : 1 hr

Learning Objectives : To learn differences blw thermodynamic Stability & Kinetic Stability.

To understand Substitution capacity Compleres.

Previous knowledge to be reminded :

Substitution → S_N1, S_N2

Bond - making & Bond - Breaking.

Topic Synopsis :

→ Each step will have an equilibrium constant. These constants are indicated as $K_1, K_2, K_3, K_4 \cdots K_n$. therefore B is equal to $B = K_1 \times K_2 \times K_3 \times K_4 \cdots \times K_n$.

→ The metal complexes generally undergo ligand exchange reactions and several other types of reactions involving the ligand.

→ The case of these type of chemical reactions [rate] is indicated by another stability constant called Kinetic stability constant.

→ Since, the kinetic stability is concerned with chemical reactions such as $ML_n + nL' \rightleftharpoons ML'_n + nL$ the activation energy of the reaction comes into play. Lower the activation energy [Ea]. Faster is the Substitution reaction

→ Based on the Case or rate of these reactions, Complexes are named as labile and inert.

Examples / Illustrations

Thermodynamic Stability of $[Co(NH_3)_6]^{+3}$ ion; $[Fe(CN)_6]^{-4}$ ion

Additional Inputs

Instantaneous, slow or moderate and very slow reactions.

Teaching Aids used

(unified) Youtube videos, PPT

References cited

unified Chemistry – Dr. O.P Agarwal
Kalayani unified chemistry – Y.R Sharma
K. Rama Rao.

Student Activity planned after teaching

Exercise with Example,
Slip test

Activity planned outside the Class Room, if any

google search on stability of different types of complexes

Any other activity

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Signature of the Lecturer

SEMESTER-I

LESSON PLAN

Name of the Lecturer: B. JOYCE Name of the Department: CHEMISTRY

Class I BBC, MBC, MPC Subject: Organic Chemistry

Name of the Topic : CHEMISTRY-[I] [INORGANIC CHEMISTRY]

Hours required : UNIT-I - Chemistry of d'-block elements. 09 hr

Learning Objectives : To study the special properties of transition metals.

Staticify of various oxidation states.

Previous knowledge to be reminded :

Classification of elements into s,p,d & f-block elements.

presence of incomplete shells.

d-block elements series.

Topic Synopsis : Definition of d-block

" when an incoming e^- enter into d-orbital it is called as d-block elements."

The General electronic configuration is $[n-1]d^{1-10}\ ns^2$

there are two incomple shells, so d-block elements are called transition elements.

There are 3 series of transition elements viz 3d, 4d & 5d series. There are 10 elements in each series.

There are 3 triads namely Cu, Ti, Cr triads in which all three elements in each series have same clue & physical properties.

Zn, Cd & Hg are called pseudo transition elements & Au, Ag & Pt are called coinage elements.

Examples / Illustrations They transition elements are crystalline, and metals with lustrous property. They are good conducters of heat and electricity.

Cu, wires good conductor of electricity.

Hg, is liquid metal

Cr & Cu Exhibit anamolous configuration

Additional Inputs

Conducters, Semi-Conducters – uses.

Teaching Aids used

Samples of conducters, Slarts with diagrams of Semi-conductors

Black-board and oral Teaching.

References cited

Text book of inorganic chemistry – Aagarwall

" " " " – cotton & wilkinson

unified chemistry – kalyani publications

Student Activity planned after teaching

Collection of Ti, Cu, Cr, Pt, Fe sample metals transition metals proparties from google search

Activity planned outside the Class Room, if any

Collection of Ti, Cu Cr, Pt, Fe sample metals.

Any other activity

→

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LESSON PLAN

Name of theLecturer: B. Joyce Name of the Department : Chemistry

Class : IMBC, NPC, BBC Subject : Organic Chemistry

Name of the Topic : Magnetic property

Hours required : 01 hour

Learning Objectives : To study magnetism

Previous knowledge to be reminded : Magnet has north & South poles opposite poles attract each other and same poles repel each other.

Topic Synopsis : Paramagnetism:- The attracting of a substance towards the magnetic field, it is called paramagnetic substance

Reason:- presence of unpaired e^-.

Ex:- Fe^{+2}, V^{+1}, Cu^{+2}

Ferromagnetism :- when magnetic intensity of a substance is increased to maximum capacity in a magnetic field. it is called ferromagnetism Reason: max. no. of e^-s.

Ex:- Cr^{+}, Mn^{+2}, Fe^{+3} etc.

Diamagnetism:- when a substance is placed in a magnetic field, if it is repelled against the magnetic field, it is called diamagnetism

Reason:- Filled e^- s

Ex:- Cu^{+}, Zn^{+2}, Cd^{+2} etc.

Magnetic moment is calculated by

Spin only formula $\mu = \sqrt{n(n+2)}$

n - no. of unpaired e^-.

Examples / Illustrations Magnetic susceptibility is determined by Cay's balence method - weight

- dd - paramagnetic
- ~~substract~~ - diamagnetic

Paramagnetic — unpaired e^-

Diamagnetic - filled e^-

Additional Inputs

Ferromagnets are used in restrictions xly ferromagnets are permanent reagent as they do not need magnetic field

Teaching Aids used

PPT presentation

References cited

unified chemistry -> Dr. O.P. Aagarwal

Student Activity planned after teaching

Memorizing the method for determining magnetic moment

Activity planned outside the Class Room, if any

Group rivison

Any other activity

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Signature of the Lecturer

LESSON PLAN

Name of the Lecturer: B. Joyce Name of the Department: Chemistry

Class: I BBC, NBC, MPC Subject: Inorganic Chemistry

Name of the Topic : Cause of lanthanide contraction

Hours required : 01 hour

Learning Objectives : To study cause of lanthanide contraction

Previous knowledge to be reminded :

The gradual decrease in the atomic and ionic size of lanthanoids with an increase in atomic number.

Topic Synopsis : Cause of lanthanide contraction :- 1. The lanthanide contraction is due to the imperfect sheilding of one 4f electron by another in the same sub shell.

2. This successive contraction accumulate and the net effect fo all the lanthanides is lanthanide contraction.

Magnetic properties :- The La^{3+} and Ce^{4+} have $4f^0$ configuration and Lu^{3+} has $4f^{14}$ configuration. These have no unpaired electron and are diamagnetic. All other f states contain unpaired and are therefore paramagnetic.

The magnetic moment of transition elements may be Calculated from the equation. $\mu_s + L = \sqrt{4S(S+1) + L(L+1)}$.

Actinides or Actinones :- The elements in which the extra electron enters 5f orbitals of $(n-2)^{th}$ main shell are known as 5f elements, actinides or Actinones. the only first four elements of this series namely Actinum, thorium, protactinum and uranium occ in nature. All the other elements are synthesised artificially bombardments. These elements are known as transuranium elem

Electronic Configuration :- $5f^{1-14}\ 6d^{1}\ 7s^{2}$

Examples / Illustrations

Cerium, Lutetium

Additional Inputs

Poor shielding of nuclear charge

Teaching Aids used

Black Board

References cited

Unified chemistry – Dr. O.P Agarwal
Kalyani Unified Chemistry – Y.R Sharma & Rama Rao.

Student Activity planned after teaching

Practice the electronic Configurations

Activity planned outside the Class Room, if any

Learning the uses of as magnetic properties — Google search. Youtube.

Any other activity

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LESSON PLAN

Name of theLecturer: B. JOYCE Name of the Department : CHEMISTRY

Class I MBC & MPC, BBC Subject : Inorg. Chemistry.

Name of the Topic : BONDING IN METALS

Hours required : 2 hrs

Learning Objectives : Metallic bond learning
Knowing the differences in types of bonding
understanding properties
of metals

Previous knowledge to be reminded :

Resonance, Conductance, electrostatic
force.

Topic Synopsis : Theories of Bonding in metals : The metal bond can be defined as "the force that binds a metal atoms to a number of electrons within its sphere of influence is known as metal bond.

Following are the main theories put forward to Explain the metallic bond.

(i) Electron sea model [Free electron theory].
(ii) valence bond theory (Pauling reasonan theory)
(iii) Band theory [Molecular orbital theory].

Electron sea model [Free electron theory].

(i) All the atoms in a metal have several unoccupied electron orbitals in their outer shells.
(ii) So, they lose some of their valency electrons and form positive ions called metal nuclei.
(iii) These free moving electrons are said to be delocalised. So the metal can be thought to be of metal cations immersed in a sea of electrons and hence this model is refferred to as an electron sea model.
(iv) electrons pull the cations towards themselves from all direvations and thus hold the cations close.

Examples / Illustrations

valence Bond theory [Pauling's Resonance theory :-
An Alternative approach of the metallic bond, on the basis of valence bond theory was developed principally by pauling. According to this theory, the bonding blw metallic atoms is covalent with resonance.

Additional Inputs

water - melon model of atom in its ionic state

Teaching Aids used

OHP & Black board

References cited

unified chemistry - Kalyani publications

unified chemistry - Jai Prakash publications

Student Activity planned after teaching

Free electron model Flexi Bands of conduction, Forbidden and non conduction in Li atom anchart.

Activity planned outside the Class Room, if any

Li element and its use as battery - Google Search

Any other activity

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LESSON PLAN

20 19 - 20 20

I, III & V - SEMS

Name of the Lecturer..... R. SHARON ROSE .....

Department..... COMMERCE .....

LESSON PLAN

Name of the Lecturer: R. Sharon Loyal Name of the Department: Commerce

Class: I M.Com (CA) Subject: CG.

Name of the Topic : Latitudes & Longitudes.

Hours required : 03.

Learning Objectives : Learn about Equator / imaginary line
Northern Hemisphere, Southern Hemisphere.
North pole, South pole.
Identification of Time, Seconds & minutes etc.

Previous knowledge to be reminded :

Mention about Equator,
Imaginary circles etc.

Topic Synopsis :

Latitudes — when earth rotates itself, it is called Rotations, & it called Axis.

The axis of North edge is called North pole &
The axis of South edge is called South pole.

Equator divides the earth into two equal parts.

The one between Equator & Northpole is called Northern Hemisphere & Southpole & Equator is called Southern Hemisphere.

Longitudes are called Meridians, which comes from Latin Word. & They are called Green which Meridians,

Development of Theories also —.

Examples / Illustrations — Calculation of Time by Longitudes
Calculation of Place by Latitudes.

Additional Inputs — Designing of Longitudes & Latitudes

Teaching Aids used — Black Board

References cited — Commercial Geography: V. Sreedhar.
Jai Bharat — By Dr. M. Sreenivasulu.

Student Activity planned after teaching — Question & Answer session

Activity planned outside the Class Room, if any — Google search about the theories of the Earth.

Any other activity —

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LESSON PLAN

Name of the Lecturer: R. Sharon Rose Name of the Department: Commerce

Class: III B.Com | V SEM Subject: CG

Name of the Topic : Indian - forestry.

Hours required : forests - features & functions.

Learning Objectives :

- To learn about the need of forests in the world

Previous knowledge to be reminded : Attention about what is forest. What are the things available in the forests. What it contains. & How it is developed.

Topic Synopsis :

Forests is a place, where living & Non-living things are exists in the place — it is a place attached to the village boundaries.

Features

1. Spacious Land.
2. Non-interference of human beings.
3. Bio-diversity.
4. Tri-peoples' existence

Functions of forests

1. Productive function.
2. Protective function.
3. Recreation function.
4. Ancillary function.

Examples / Illustrations

2. Examples of Indian forests.
Nallamala forests.
Rompachoodavaram forests } mentioned

Additional Inputs

Teaching Aids used — Powerpoint Presentation.

References cited — Kalyani Publications, by D.R. Khullar

Student Activity planned after teaching — Slip Test.

Activity planned outside the Class Room, if any — given Seminar Topic on Types of forests.

Any other activity —

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Signature of the Lecturer

LESSON PLAN

Name of the Lecturer: R J Lovon Rose   Name of the Department: Commerce

Class: III I Com (R)   Subject: CG

Name of the Topic : Rights & Acts of forests.

Hours required : 03.

Learning Objectives :

To learn about what are the rights of tribal people.

Previous knowledge to be reminded :

Mention about How to improve forests, what is Afforestation & what is deforestation etc.

Topic Synopsis :

Forest Rights Act - 2006 :-

This act is also called Girijans & Conv- enforced Forest People's living Act. This Acts is meant for welfare of tribals, people lives in hill area etc.

This Act provides three types of rights.

1. Land holding Rights. 2. Consumption Right
3. Forest Convervation Right.

Chipko Movement. & Forest Convervation Act Compensatory Afforestation Fund Management Authority.

Measures to protect Earth :-

1. Urban Forest Programme.
2. Agricultural Forest Programme.

Examples / Illustrations

Example : Chipko Movement.

Additional Inputs — Various Acts related to Forest Rights are mentioned.

Teaching Aids used — Black Board.

References cited — Kalyani Publishers

Student Activity planned after teaching — Question & Answers.

Activity planned outside the Class Room, if any — Google search for ~~Forest~~ rights Act.

Any other activity —

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Signature of the Lecturer

LESSON PLAN

Name of the Lecturer: R. Sharon Raj. C. Name of the Department: Comm

Class: III B.Com / V Section Subject: CFA

Name of the Topic : Cost Accounting - Meaning

Hours required : 04.

Learning Objectives :

To learn about Meaning of Cost Accounting. Limitations of Cost Accounting. objectives of Cost Accounting etc.

Previous knowledge to be reminded :

Mention about various Types of costing, Like uniform Costing, Marginal Costing, Standard costing, Historical Costing, Direct Costing, & Absorption Costing.

Topic Synopsis :

⇒ features, objectives & functions of Cost Accounting.

⇒ Scope, Advantages & Limitations of Cost Accounting.

⇒ Relationship between Cost, Management & financial accounting.

Cost sheet & statements of Cost.

Calculation of Prime cost, Works Cost, Cost of production & Calculation of Total Cost & Cost of Sales.

Examples / Illustrations — Problem taken as Calculation of Prime Cost, Works Cost, & other costs & finally Calculation of Cost of Sales

Additional Inputs — Practicing Problems in Text Book.

Teaching Aids used — Black Board

References cited — Kalyani Publishers by S.P. Jain. & K.L. Narang

Student Activity planned after teaching → Question & Answer Session

Activity planned outside the Class Room, if any — Slip Test.

Any other activity —

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LESSON PLAN

Name of theLecturer: R. Sharon Raje Name of the Department: Commerce

Class: III B.Com (Gen & Bus) Subject: Cost Acc

Name of the Topic : Methods of Pricing Issues.

Hours required : 10.

Learning Objectives :

To study about the Issue of Materials.

Previous knowledge to be reminded :

Mention about various Levels of materials & to control the inventory.

Topic Synopsis :

→ FIFO — First-in-first out.

→ LIFO — Last in first out.

→ Weighted average Method.

→ Simple Average Method.

→ Base stock Method with FIFO & LIFO.

→ Market Price Method.

Examples / Illustrations

→ Problems & Illustrations in FIFO, LIFO, Simple/weighted Average, Base stock Method.

Additional Inputs

Problems on LIFO
FIFO
Weighted average method given.

Teaching Aids used

Black Board.

References cited

— Kalyani — Publishers, Technical Publications

Student Activity planned after teaching

— Slip Test
Assignment.

Activity planned outside the Class Room, if any

Doing Problems.

Any other activity

—

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LESSON PLAN

2018 - 2019

Name of the Lecturer..D. V. Ramana.....

Department..Computer Science.....

LESSON PLAN

Name of theLecturer: D.V. Ramana ..... Name of the Department Computer Science

Name of the Topic : Number system : Binary, hexa & octa numbering system.

Hours required : 2 hrs

Learning Objectives :

* working with Binary numbers.

Previous knowledge to be reminded :

* Computers are electronic machines that operate using binary logic

Topic Synopsis :

* converting a Binary number to Decimal form.

- In binary no:, all the colomns are powers of 2.

- converting a Hexa decimal no. to Binary Form.

* Octal Number System :-

	Fourth Digit	Third Digit	Second digit	First digit.
Decimal	10^3	10^2	10^1	10^0
Binary	2^3	2^2	2^1	2^0
octal	8^3	8^2	8^1	8^0.

Examples / Illustrations

Ex:- 1101 into decimal no.

Decimal number $= 1 \times 2^3 + 1 \times 2^2 + 0 \times 2^1 + 1 \times 2^0$

$= 1 \times 8 \times 1 \times 4 + 0 \times 2 + 1 \times 1$

$= 8 + 4 + 0 + 1$

$= 13$.

Additional Inputs

Teaching Aids used

Black board.

References cited

* Referenced by REEMA THAREJA.

Student Activity planned after teaching

* converting on Binary number system.

Activity planned outside the Class Room, if any

Any other activity

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Ramana.D.V
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: D.V.Ramana ..... Name of the Department : Computer Science

Name of the Topic : Types of softwares

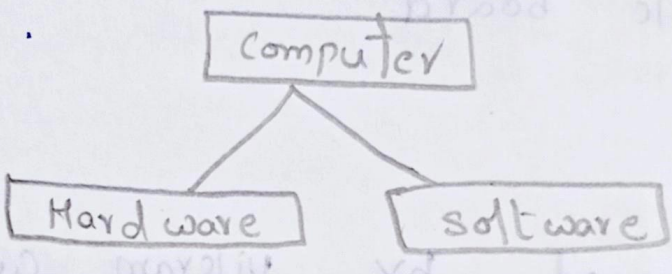
Hours required : 2 hrs

Learning Objectives :

* computer software
* computer Hardware

Previous knowledge to be reminded :

* Basically computer system is made up of these two terms.



Topic Synopsis :

* computer H/w is nothing but collection of actual physical components of system. so the input, storage, processing, C.P.U, output devices are hardware.
* software is a program.
* program is a set of instructions.
* A computer software can be broadly classified into four categories;

1. system software
2. Application software
3. Firm ware
4. Live ware.

Examples / Illustrations

Assembly language program → [Assembler] → Machine language Program.

High level language program → [compiler] → Machine language Program.

High level language program → [Interpreter] → Machine language program.

Additional Inputs

Teaching Aids used

Black board

References cited

*Referenced by vikram Gupta.

Student Activity planned after teaching

*Assignment of software & its types.

Activity planned outside the Class Room, if any ---

Any other activity ---

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Ramana.D.V
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: D.V. Ramana ..... Name of the Department : Computer science

Name of the Topic : Page layout & Back grounds in photoshop

Hours required : 1hr

Learning Objectives : * width & Height
* color Modes.
* Background

Previous knowledge to be reminded :

Topic Synopsis : *

Pixels
✓Inches.
centimeters.
Milli " .
Points
Picas
colomns.

* If you did not select & choose to create a custom size, then this is where you can type in your desired document width & Height.

* color Modes :-

Bitmap
Grayscale.
✓RGB color
Lab color.

* color mode allows you to choose what mode & bit you wish to use for your new document.

* Back ground contents :-

✓white
Back ground color
Transparent.

* This allows you to determine the color of your background layer choosing white.

Examples / Illustrations ——

Additional Inputs ——

Teaching Aids used

ppts. & Black board.

References cited

* Referenced by V.S.N. Kumar.

Student Activity planned after teaching

* create your visiting card.
* Design a passport photo.

Activity planned outside the Class Room, if any ——

Any other activity ——

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Ramana. D V
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: D.V.Ramana ..... Name of the Department : Computer science

Name of the Topic : working with tool Box :- Pen, Erasers, brushes

Hours required : 1hr.

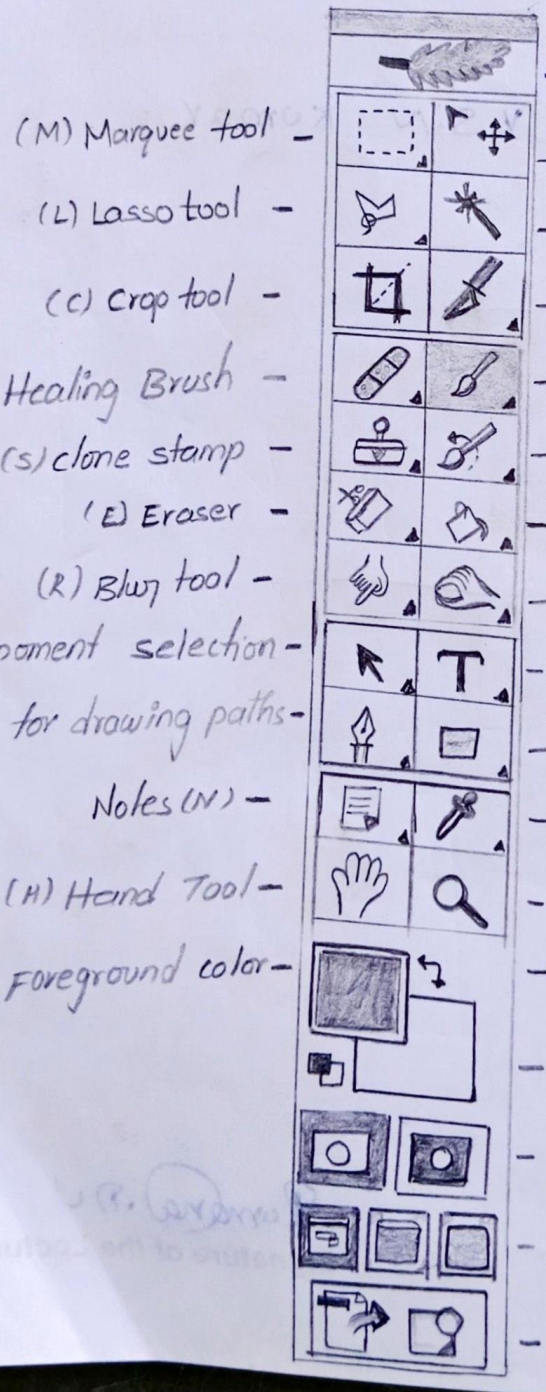
Learning Objectives :

* Tool Box

Previous knowledge to be reminded :

* Some tools in the Tools panel have options that appear in the context sensitive option bar.

Topic Synopsis :



Examples / Illustrations ———

Additional Inputs ———

Teaching Aids used

* PPTs | Black board

References cited

* Referenced by V.S.N. Kumar

Student Activity planned after teaching

* convert color photo to black & white photo.
* Background changes.

Activity planned outside the Class Room, if any ———

Any other activity ———

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Ramana. DV
Signature of the Lecturer

LESSON PLAN

Name of theLecturer: D.V.Ramana ..... Name of the Department : Computer Science

Name of the Topic : How to create addy, noice, brush, light effects, filter, printing etc

Hours required : 8 hrs.

Learning Objectives :

*Noise filters.

Previous knowledge to be reminded :

- know the filter Menu.

Topic Synopsis : * The noise filters add (or) remove noise, (or) pixels with randomly distributed color levels.

* This helps to blend a selection into the surrounding pixels.

* Noise filters can create unusual textures (or) remove problem areas, such as dust & scratches.

* To reduces noise while preserving edges based on user settings affecting the overall image (or) individual channel.

* clumps pixels into a solid color in a polygon shape.

Examples / Illustrations —

Additional Inputs —

Giving the notes.

Teaching Aids used

ppt. | Black board.

References cited

Referenced by V.S.N kumar.

Student Activity planned after teaching

Practice the Filters & adds

Activity planned outside the Class Room, if any →

Any other activity —

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Ramana .D

Signature of the Lecturer